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Guillaume Favre

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$o_2 + 2 nacl \rightarrow 2 naf + cl_2$ $2 h_2 + 1 o_2 \rightarrow 2 h_2o$ $pb + 2 hcl \rightarrow 2 h_2o + 1 pbcl_2$
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calcium carbonate is heated and
decomposes to solid calcium oxide and
carbon dioxide gas gaseous butane $c_4 h_{10}$ reacts with diatomic oxygen gas to
yield gaseous carbon dioxide and water
vapor

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equation is an idea the concept of
telling the chemist that if the atoms
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there are balanced amounts of the
atoms molecules and compounds on the
other side in the product

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equation is to write down the chemical

formula of reactants that are listed on

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after this you can list down the

products on the right hand side of the

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6 kbr 1 al 2 so 4 3 7 1 ch 4 2 o 2 1 co 2

2 h 2 o 8 1 c 3 h 8 5 o 2 3 co 2 4

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atoms participating in a reaction the reactants products and direction of the reaction balancing an unbalanced equation is mostly a matter of making certain mass and charge are balanced on the reactants and products side of the reaction arrow

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