

# Oh But My Darling What If You Fly

## L Darling-Hammond

**Question #44ec4 - Socratic** That's a silly way to say carbon dioxide!  $\text{CO}_2(\text{g}) + \text{Ca}(\text{OH})_2(\text{aq}) \rightarrow \text{Ca}^{2+}(\text{aq}) + \text{H}_2\text{CO}_3(\text{aq})$  May be wrong, but it seems reasonable. The gas would interact with the hydroxide and water molecules to form carbonic acid. There are also concentrations of  $\text{OH}^-$  and  $\text{H}^+$  leftover from this reaction and from the auto-ionization of water, but that's probably going too far into detail. **Question #627ef - Socratic** functional groups like -OH, -CHO, -CO- are present in carbohydrates. commonly two functional are present in a carbohydrate, -OH and -CHO, or -OH and -CO-. Aldose shows different reactions and test like Tollen's silver mirror test, felling's test. different types of reactions are also shown by carbohydrates like Osazone synthesis, length increasing reaction, ... **What are the spectator ions in the reaction between sodium** Oct 16, 2016 · The nitrate and the natrium ions.  $\text{Na}_2\text{CO}_3(\text{aq}) + 2\text{AgNO}_3(\text{aq}) \rightarrow \text{Ag}_2\text{CO}_3(\text{s}) + 2\text{NaNO}_3(\text{aq})$  The net ionic equation is:  $2\text{Ag}^+(\text{aq}) + \text{CO}_3^{2-}(\text{aq}) \rightarrow \text{Ag}_2\text{CO}_3(\text{s})$  ... **Could a buffered solution be made by mixing aqueous ... - Socratic** Mar 25, 2017 · Could a buffered solution be made by mixing aqueous solutions of HCl and NaOH? Explain. Why isn't a mixture of a strong acid and its conjugate base considered a buffered solution? **What is the product of the following reaction? 1)CH<sub>3</sub>OH - Socratic** Dec 23, 2017 · These are ostensibly acid-base reactions.... For ammonium we could write....  $\text{NH}_4^+ + \text{HO}^- \rightarrow \text{NH}_3(\text{aq}) + \text{H}_2\text{O}(\text{l})$  For methanol, the acid base reaction would proceed ... **0.10 moles of sodium cyanate dissolved in 250cm<sup>3</sup> distilled** Dec 28, 2015 · The cyanate anions will react with water to form isocyanic acid, HNCO, and hydroxide anions,  $\text{OH}^-$ . This tells you that you can expect the pH of the resulting solution to be higher than 7. Calculate the molarity of the sodium cyanate solution by using the given number of moles and volume. Keep in mind that you need to use liters of solution and ... **What is the formula of "magnesium hydroxide"?** | **Socratic** Feb 27, 2017 ·  $\text{Mg}(\text{OH})_2$  Now we know that hydroxides are salts of  $\text{HO}^-$ , and some metal ion. Now if the parent metal has an electronic configuration of 2:8:2, then there are 12 electrons, and the atomic number of the metal is equal to 12. We look on the Periodic Table, and we find that  $Z=12$ , for "magnesium metal". As a Group 2 metal, magnesium forms a  $\text{Mg}^{2+}$  ion, and hence its hydroxide is  $\text{Mg}(\text{OH})_2$ . **What is the mass percent of hydrogen for C<sub>3</sub>H<sub>5</sub>OH?** | **Socratic** Jan 2, 2017 · In order to compute the mass percent of hydrogen in  $\text{C}_3\text{H}_5\text{OH}$  we're going to need two things: 1: The formula mass of the entire compound 2: The atomic mass of hydrogen The formula mass of  $\text{C}_3\text{H}_5\text{OH}$  is 58.08 g mol because:  $3(12.01 \text{ g mol}) + 6(1.008 \text{ g mol}) + (16.00 \text{ g mol}) = 58.08 \text{ g mol}$  The atomic mass of the hydrogen atom is 1.008 g mol Now,

we have to use the following equation: The numerator ... **What is the chemical formula for lithium hydroxide? - Socratic** Apr 15, 2017 · Lithium is a Group 1 metal and commonly forms a  $M^+$  ion. Hydroxide anion,  $-OH$ , has a unit negative charge. When they make music together, there is thus 1:1 stoichiometry between ions:  $Li(s) + H_2O(l) \rightarrow LiOH(aq)$   
 $+ 1/2 H_2(g) \uparrow$  **If the pH=9 for a solution of KOH(aq), what is the ... - Socratic** Aug 18, 2016 · If the  $pH = 9$  for a solution of  $KOH(aq)$ , what is the concentration of  $H^+$ ?

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