

Le Coureur Du Campus

Daniel F McAuley

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web limiting reactant theoretical and percentage yields key 4 $\text{ko}_2 \text{ s } 2 \text{ h } 2 \text{ o l } 4 \text{ koh aq } 3 \text{ o}_2 \text{ g}$ if a reaction vessel contains 0 mol ko 2 and 0 mol h 2 o what is the limiting reactant

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web limiting reactant reactant present in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage of the theoretical yield

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$$\text{Cu} + 2\text{AgNO}_3 \rightarrow \text{Cu(NO}_3)_2 + 2\text{Ag}$$

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$$3\text{CO} + 7\text{H}_2 \rightarrow \text{C}_3\text{H}_8 + 3\text{H}_2\text{O}$$

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web start with 8 g mole h₂s and multiply it with 2 mole h₂o over 2 mole h₂s from the equation

$$8 \text{ g mole h}_2\text{s} \times \frac{2 \text{ mole h}_2\text{o}}{2 \text{ mole h}_2\text{s}} = 18 \text{ g h}_2\text{o}$$

149 g h₂o is your theoretical yield so percent yield

$$\frac{137 \text{ g h}_2\text{o}}{149 \text{ g h}_2\text{o}} \times 100 = 91.77\%$$

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$$\text{C}_3\text{H}_6 + 3\text{O}_2 \rightarrow 3\text{CO} + 3\text{H}_2\text{O}$$

if the reaction of 21 g of o₂ produces a 47 g yield how many grams of co would be

produced step one figure out how many moles of the limiting reagent you have

$$\frac{21 \text{ g}}{32 \text{ g/mol}} = 0.66 \text{ mol}$$

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web percent yield is a measure of the efficiency of the experimental design yield efficiency mass of product obtained calculated mass of the product expected x100 in this reaction a yield 80 is anticipated

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web a write a balanced equation for the reaction

$$4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$$

b determine the theoretical yield of al₂o₃

$$160 \text{ g Al} \times \frac{1 \text{ mol Al}}{27 \text{ g Al}} \times \frac{2 \text{ mol Al}_2\text{O}_3}{4 \text{ mol Al}} \times \frac{101.96 \text{ g Al}_2\text{O}_3}{1 \text{ mol Al}_2\text{O}_3} = 302.3 \text{ g}$$

c determine the percent yield

$$\frac{\text{actual}}{\text{theoretical}} \times 100 = \frac{260 \text{ g}}{302.3 \text{ g}} \times 100 = 86.01\%$$

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web 7 na 2 c 2 o 4 is the limiting reactant percent yield 86.6% only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 c is the limiting reagent 433 g of h₂ are left over 13

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