

Plinio El Viejo Bebida

Michael Brown

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object s vertical lines are drawn

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imagine that you have an object suspended by transparent threads

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figure 7 an isometric view with dimensions we have dimensioned the object in the isometric drawing in figure 7 as a general guideline to dimensioning try to think that you would make an object and dimension it in the see more

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identification of

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given that the boiling points of ethanol and methanol 78 c and 64 c respectively are far off in a significant way this seems like a valid conclusion conclusion thermometer calibrations are necessary in order to collect reliable information as precise calculations are imperative to gathering correct results

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melting point 1 a push the open end of a capillary tube into the powdered acetamide b move the powder to the closed end of the capillary tube by tapping it on the table repeat until the the powdered acetamide occupies 1 2 mm of the capillary tube end c

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based on its boiling point background information objekte can be distinguished into acid properties and physical properties

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discussion and conclusion the technique recrystallization was used because it is a simple method to purify a solid compound the percentage yield for benzoic acid and acetanilide acid being high indicated the initial sample had a low concentration of impurities with the melting point being in a specific

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introduction the boiling point of a

compound is the temperature at which it changes from a liquid to a gas this is a physical property often used to identify substances or to check the purity of the compound it

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conclusion after completing this experiment the melting point of three different unknown compounds was obtained using the mel temp apparatus compound a had a melting point range of 114 6 115 5 c thus revealing it to be acetanilide compound b had a melting point range of 122 6 124 5 c thus
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 melting point can be used as the identity and purity of an organic solid however it is not the same as the boiling point the boiling point is a characteristic unique of pure liquids the boiling point is the temperature at which the vapor pressure of a liquid is equal to the atmospheric pressure

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the purpose of the experiment is to understand what vapour pressure is and how it changes with temperature to understand the relationship between vapour pressure at a given temperature and boiling temperature at a given

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50 c so in order to convert an experimental measurement taken in calgary which is at higher altitude than sea level and so is at a lower pressure to that reported for sea level higher pressure one needs to add a correction factor since the boiling point at sea level is higher than that at higher altitudes

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a sample's boiling point can be

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and by using a thiele tube the most
straightforward method uses a thiele

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