

Dachau Arbeit Macht Frei

Daniel F McAuley

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masses of calorimeter th water in it and the hot object made
up of copper which is put in the calorimeter are the same the
initial temperature of the calorimeter and water is 30 c and
that of hot object is 60 c the specific heat of copper and
water are 0 09 cal gm c and 1 cal gm c respectively

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web oct 28 2019 solution use this equation $q = mc\Delta t$ where q is heat flow m is mass in grams and Δt is the
temperature change plugging in the values given in the
problem you get $q_{\text{water}} = 418 \text{ J/g} \times 110 \text{ g} \times 26.6^\circ\text{C}$ $q_{\text{water}} = 550 \text{ J}$ $\Delta h_{\text{water}} = 550 \text{ J}$

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 t q heat m mass q_{rxn} q_{sur} t t_f t_i name per date c specific
heat for water 4184 J/g°C 1 what is the specific heat of
aluminum if the temperature of a 28.4 g sample of aluminum
is increased by 8.1 °C when 207 J of heat is added 2

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released or absorbed at constant pressure is equal to Δh the
relationship between heat and Δh_{rxn} is $\Delta h_{\text{rxn}} = q_{\text{rxn}}$
 $q_{\text{calorimeter}} = -m c \Delta t$ $7.3.14.7.3.14. \Delta h_{\text{rxn}} = q_{\text{rxn}} = -q_{\text{calorimeter}}$
 $m c \Delta t$

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web use the information in this chart to answer the following
questions 1 if i have 125 grams of water at a temperature of
25.0 °C and heat it to 145.0 °C what will the change in enthalpy
be 2 sketch the heating curve temperature vs time for the
process in 1 3 i have 18.7 grams of an unknown liquid
chemical at a temperature of 35.0 °C

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experiment 25.0 mL of 1.00 M HCl at 25.0 °C is added to 25.0
mL of 1.00 M NaOH at 25.0 °C in a foam cup calorimeter a
reaction occurs and the temperature rises to 32.0 °C calculate
the enthalpy change in kJ for this reaction assume the
densities of the solutions are 1.00 g/mL and that

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chemical reactions consider a simple example that illustrates
the core idea behind calorimetry suppose we initially have a
high temperature substance such as a hot piece of metal m
and a low temperature substance such as cool water w

answer the initial temperature of the copper

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web choose 1 answer the final temperature is closer to t_1 than to t_2 a the final temperature is closer to t_1 than to t_2 the final temperature is exactly halfway between t_1 and t_2 b the final temperature is exactly halfway between t_1 and t_2 the final temperature is closer to t_2 than to t_1 c

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web 6 a 218 g sample of steam at 121 $^{\circ}\text{C}$ is cooled to ice at 14 $^{\circ}\text{C}$ find the change in heat content of the system 7 in going from ice at 34 $^{\circ}\text{C}$ to steam at 138 $^{\circ}\text{C}$ a sample of water absorbs $1 41 \times 10^5 \text{ J}$ find the mass of the sample answers 1 2 68 x 104 J 2 6 23 x 104 J 3 1 71 x 104 J 4 1 11 x 105 J 5 2 11 x 105 J 6

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