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fusion 9.37 kJ/mol how much energy is

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which will likely have the larger

magnitude ΔH fusion or ΔH

vaporization explain your reasoning

answer a

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faster and faster as they absorb kinetic energy

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web if we are dealing with a gas we still use two equations $q = mc\Delta T$ would be the specific heat of gases water and then here it would be the specific heat of ice

now again going back to the question we are looking for liquid water so that would mean the answer is option c line segment cd would have the specific heat of liquid water

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be the same heating is when energy is transferred to the substance and the kinetic energy of the molecules increases red arrows to the right cooling is when energy is transferred away from the substance dissipated to the surroundings and the kinetic energy of the molecules decreases blue arrows to the left

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