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each would be written as

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alcohol n_b 23 g 46 g mol 1 0 5 mol number of moles of
solvent water n_a 54 g 18 g mol 1 3 mol total number of
moles $n_a + n_b$ 0 5 3 3 5 mol mole fraction of solute ethyl
alcohol $x_b = \frac{n_b}{n_a + n_b} = \frac{0.5}{3.5} = 0.1429$

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alcohol in 40 g of H_2O molar mass of H_2O is 18 gm mole and
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4.5 32 0.14 mole moles of H_2O 40 18 2.2 moles therefore
according to the equation

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mixtures where the partial pressure of a gas was its mole

fraction time the total pressure $p_a + p_t$ mole fraction a left
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mol b note in the given mixture the sum of all the mole
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fractions are multiplied by 100 they give the mole
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