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melting point of 23.2°C and has a ΔH
fusion 9.37 kJ/mol how much energy is
required to melt 263.1 g TiCl_4 for TiCl_4
which will likely have the larger
magnitude ΔH fusion or ΔH
vaporization explain your reasoning
answer a

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graphs they plot a substance s
temperature y axis against heat x axis
for heating curves we start with a solid
and add heat energy for cooling curves
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multiple choice 2 minutes 1 pt describe

the substance between letters d and e
 gas liquid melting evaporating 2
 multiple choice 2 minutes 1 pt between
 which points is the temperature of the
 substance remaining constant a b only
 a b c d e f b c only b c d e 3 multiple
 choice 2 minutes 1 pt

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web this plot of temperature shows
 what happens to a 75 g sample of
 steam initially at 1 atm and 200 c as
 heat is removed at a constant rate a b
 cooling steam b c condensing steam c d
 cooling liquid water to give a
 supercooled liquid d e warming the
 liquid as it begins to freeze e f freezing
 liquid water f g cooling ice

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curve worksheet if this curve is read
 from right to left it is a cooling curve
 the diagram below illustrates the steps
 involved to convert 10g of solid ice at
 20 c to 10g of gaseous steam at 140 c a
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 construction and use of a typical phase
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 imagine that you have a block of ice
 that is at a temperature of 30 text o
 text c well below its meltingpoint the
 ice is in a closed container as heat is
 steadily added to the ice block the
 water molecules will begin to vibrate
 faster and faster as they absorb kinetic
 energy

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web stearic acid c17h35cooh is cooled
 at a constant rate the temperature of
 the sample is recorded at 2 minute
 intervals in the data table below read

more heating cooling curves heating cooling curves page 1 heating cooling curves a the following graph is a heating curve showing the addition of heat at a constant rate of 500 0 joules minute

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web if we re dealing with a gas we d still use two equals n cac would be the specific heat of gasses water and then here it would be the specific heat of ice now again going back to the question we re looking for liquid water so that would mean the answer is option c line segment cd would have the specific heat of liquid water

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web 1 the heating curve below

represents a sample of a substance starting as a solid below its melting point and being heated over a period of time which statement describes the energy of the particles in this sample during interval de a both potential energy and average kinetic energy increase

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web heating and cooling graphs tend to be the same heating is when energy is transferred to the substance and the kinetic energy of the molecules increases red arrows to the right cooling is when energy is transferred away from the substance dissipated to the surroundings and the kinetic energy of the molecules decreases blue

arrows to the left

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web questions 5 marks for this experiment identify the independent variable dependent variable control variable mark mark mark name the process in which matter change from solid to liquid 1 name the process in which matter change from liquid to solid 1

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