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Didier Musso

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therefore the reaction is exothermic or heat releasing 2

answer a 1 explanation the entire reaction was multiplied by

2 therefore Δh must also be halved 3 answer c explanation

option c is incorrect because 566 kJ mol is the Δh per mole of O_2 and not for 2 moles

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both mass and enthalpy changes $H_2O \rightarrow H_2O(l)$ 6 01 kJ mol

this is an endothermic process it requires 6 01 kJ to melt one

mole of ice $H_2O(s)$ the enthalpy value will change if the

number of moles varies from the 1 1 reaction stoichiometry

this is an exothermic process

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main concepts of thermochemistry such as the relationship

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calorimetry constant volume calorimetry the enthalpy the

standard enthalpies of formation and their use in

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web we will revisit this today and go on to see that if we use

a special kind of thermochemical reaction called the

standard enthalpy of formation we can calculate enthalpies

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individual thermochemical equations for each step

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web we can express the first law in many ways one of the

more useful expressions is that the change in internal energy

ΔU of a system in any process is equal to the heat q added to

the system plus the work w done on the system by its

surroundings $\Delta U = q + w$

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