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 combining a strong acid and a strong
 base will produce salt and water excess
 strong acid or strong base will not act
 as a buffer
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 that is they do not ionize fully in
 aqueous solution a solution of a weak
 acid in water is a mixture of the
 nonionized acid hydronium ion and the
 conjugate base of the acid with the
 nonionized acid present in the greatest
 concentration
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 salts such as cacl 2 are neutral because
 it is created from hydrochloric acid a
 strong acid and calcium hydroxide a
 strong base an aqueous solution of
 nanh 2 is basic because it can
 deprotonate alkynes alcohols and a host
 of other functional groups with acidic
 protons such as esters and ketones
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