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form of finely split powder copper
metal is produced as a byproduct of the
process as iron ions the iron metal
dissolves

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understanding of stoichiometry a
reaction between iron and copper ii
sulfate solution will be conducted this
will help you to differentiate limiting
and excess reactant in a chemical
reaction finally the theoretical and
percent yield of this reaction will be
calculated

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you will also use stoichiometry to
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clean dry 100 ml beaker and record 2
measure out 8 00 grams of copper ii
sulfate onto a piece of filter paper
taking the filter paper's mass into
account add these to the beaker 3
measure 50 0 ml of water in the
cylinder and add it to the beaker 4
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web through this experiment we sought to use stoichiometric principles which derive from stoichiometry the study of substance quantities in ratio or proportion in a chemical reaction in this lab experiment the principles help distinguish the equation between iron

powder and aqueous copper ii sulfate in the iron copper ii sulfate redox reaction

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web this experiment is focused on a reaction between iron metal and a solution of copper ii sulfate we are able to figure out the chemical equation between the two by using stoichiometric principles the iron displaces the copper ii sulfate ions the copper ii is then reduced and gains electrons

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web stoichiometry can be defined as the relationship between the relative quantities of substances taking part in a reaction or forming a compound stoichiometric principles can be used to deduce the appropriate chemical equation for a reaction in our experiment we used solid iron in the form of powder was added to a solution of copper ii sulfate

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aqueous copper ii sulfate in the iron copper ii sulfate redox reaction the iron ions dissolve in the copper ii sulfate solution to produce copper ii ions

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web under the finalize of the experiment it was found that iron was the limiting reactant and iron ii sulfate was a product of the reaction between 7 0002g away cooper ii sulfate and 2 0101g iron metal which produced 2 3037g of copper product with a percent yield of 100 72 introduction

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period 5 in this experiment i observed the reaction of iron with copper ii sulfate i obtained the results to this experiment through adding heated CuSO_4 solution to the iron powder stirring the solution adding water then removing water adding acetone heating the acetone on low

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