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## N Noddings

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 web limiting reactant theoretical and percentage yields key 4  $\text{K}_2\text{S} + 2 \text{H}_2\text{O} + 4 \text{KOH} \rightarrow 3 \text{O}_2 + \text{G}$  if a reaction vessel contains 0 mol  $\text{K}_2\text{O}$  and 0 mol  $\text{H}_2\text{O}$  what is the limiting reactant  
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### section 12 3 limiting reagent

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limiting reagent you have 0.216 moles 32 grams/mole = 6.912 grams 0.66 moles 32 g/mole = 21.12 grams

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web a write a balanced equation for the reaction  $\text{Al} + \text{O}_2 \rightarrow \text{Al}_2\text{O}_3$  b determine the theoretical yield of Al<sub>2</sub>O<sub>3</sub> 1600 g Al  $\times \frac{1\text{ mol Al}_2\text{O}_3}{2\text{ mol Al}} = 101.96\text{ g Al}_2\text{O}_3$  3023 g Al<sub>2</sub>O<sub>3</sub> 26982 g Al 4 mol Al 1 mol Al<sub>2</sub>O<sub>3</sub> c determine the percent yield  $\frac{\text{yield actual}}{\text{yield theoretical}} \times 100 = \frac{2600\text{ g}}{3023\text{ g}} \times 100 = 86.01\%$

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web 7 Na<sub>2</sub>C<sub>2</sub>O<sub>4</sub> is the limiting reactant percent yield 86.68 only four molecules can be made 9 this amount cannot be weighted by ordinary balances and is worthless 10 nitrogen is the limiting reagent 11 yes methane is the limiting reagent 12 C is the limiting

reagent 4 33 g of h<sub>2</sub> are left over 13  
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web limiting reagents and percentage yield worksheet 1 consider the reaction  $\text{I}_2\text{O}_5 + \text{CO} \rightarrow \text{CO}_2 + \text{I}_2$  80 0 grams of iodine v oxide i<sub>2</sub>o<sub>5</sub> reacts with 28 0 grams of carbon monoxide co determine the mass of iodine i<sub>2</sub> which could be produced 80 g i<sub>2</sub>o<sub>5</sub> 1 mol i<sub>2</sub>o<sub>5</sub> 1 mol i<sub>2</sub> x s 1 333 8 g i<sub>2</sub>o<sub>5</sub> 1 mol i<sub>2</sub>o<sub>5</sub> 28 g co 1 mol co

**7 3 limiting reactant and percent yield problems** - Jun 13 2023

web jun 2 2020 the percent yield of a reaction is the ratio of the actual yield to the theoretical yield multiplied by 100 to give a percentage text percent yield text actual yield g over text theoretical yield g times 100 label 3 7 3