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 with 0 100 m sodium hydroxide the values of the ph  
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 naoh are listed in the first column of this table and are  
 graphed in

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 neutralisation reaction between an acid and a base can be  
 very useful if an acidic solution of known concentration a  
 standard solution is added to a basic alkaline solution of  
 unknown concentration until the solution is exactly  
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web the above equation works only for neutralizations in  
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 hydroxide

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 analysis method used to determine the concentration of an  
 acid or base by neutralizing the acid or base with a known  
 concentration standard solution the concentration of a  
 solution can be determined by knowing the acid and base  
 dissociation constant

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 and procedures however the most common types of titration  
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of a known acidic or basic substance through acid base reactions the analyte titrand is the solution with an unknown molarity the reagent titrant is the solution with a known molarity that will react with the analyte

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web as seen in the chapter on the stoichiometry of chemical reactions titrations can be used to quantitatively analyze solutions for their acid or base concentrations in this section we will explore the underlying chemical equilibria that make acid base titrimetry a

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