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in an amount lower than required by the reaction stoichiometry thus limiting the amount of product generated percent yield measure of the efficiency of a reaction expressed as a percentage of the theoretical yield

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web start with 8 3 mole H_2S and multiply it with 2 mole H_2O over 2 mole H_2S from the equation $8 3 \text{ mole H}_2\text{S} \times 2$

$\text{mole H}_2\text{O} = 2 \text{ mole H}_2\text{S} \times 18 \text{ g H}_2\text{O} / 1 \text{ mole H}_2\text{O} = 149 4 \text{ g H}_2\text{O}$ is your theoretical yield so percent yield $137 1 \text{ g H}_2\text{O} / 149 4 \text{ g H}_2\text{O} \times 100 = 91 77\%$ problems 6 8 are similar to problem 5

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web limiting reagents problem one 1 for the balanced equation $\text{C}_3\text{H}_6 + 3 \text{O}_2 \rightarrow 3 \text{CO} + 3 \text{H}_2\text{O}$ if the reaction of 21 3 grams of O_2 produces a 47 3 yield how many grams of co would be produced step one figure out how many moles of the limiting reagent you have $0 216 / 32 = 6 75 \text{ moles}$ 32 g m

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web percent yield is a measure of the

efficiency of the experimental design
 yield efficiency mass of product
 obtained calculated mass of the product
 expected $\times 100$ in this reaction a yield
 80 is anticipated

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web a write a balanced equation for the
 reaction $\text{Al}_2\text{O}_3 + 4\text{Al} + 3\text{O}_2 \rightarrow 2\text{Al}_2\text{O}_3$ b
 determine the theoretical yield of Al_2O_3
 $160.0\text{ g Al} \times \frac{1\text{ mol Al}}{27.0\text{ g Al}} \times 2\text{ mol Al}_2\text{O}_3 \times$
 $\frac{101.96\text{ g Al}_2\text{O}_3}{302.3\text{ g Al}_2\text{O}_3} = 26.982\text{ g}$
 Al_2O_3 c determine the
 percent yield $\frac{\text{yield actual}}{\text{yield theoretical}} \times 100$
 $\frac{26.0\text{ g}}{302.3\text{ g}} \times 100 = 86.01\%$ theoretical

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 web 7 na 2 c 2 o 4 is the limiting
 reactant percent yield 86.6% only four
 molecules can be made 9 this amount
 cannot be weighted by ordinary
 balances and is worthless 10 nitrogen is
 the limiting reagent 11 yes methane is
 the limiting reagent 12 c is the limiting
 reagent 4 33 g of H_2 are left over 13
 6 5 limiting reagent and percent yield
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 determine theoretical yields actual
 yields and percent yields identify a

limiting reagent from a set of reactants
 calculate how much product will be
 produced from the limiting reagent
 calculate how much reactant s remains
 when the reaction is complete

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 web the percent yield is calculated as
 follows $\text{percent yield} = \frac{\text{actual yield}}{\text{theoretical yield}} \times 100$
 based on this
 definition we would expect a percent
 yield

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 identify the limiting reactant limiting
 reagent in a given chemical reaction
 calculate how much product will be
 produced from the limiting reactant
 calculate how much reactant s remains
 when the reaction is complete
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methanol CH_3OH can be produced by the
 following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$
 calculate the theoretical yield of CH_3OH
 if 68.5 g of CO is reacted with 8.6 g of
 H_2

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 excess calculation name sto 4 solve
 stoichiometric problems from a
 balanced chemical equation 3 sto 5
 differentiate between the actual yield
 and theoretical yield of a chemical
 reaction

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 practice name 1 consider the following
 reaction $\text{NH}_4\text{NO}_3 + \text{Na}_3\text{PO}_4 \rightarrow \text{NH}_4\text{PO}_4$
 NaNO_3 which reactant is limiting
 assuming we started with 30.0 grams of
 ammonium nitrate and 50.0 grams of
 sodium phosphate what is the
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 yield worksheet 1 consider the reaction
 $\text{I}_2\text{O}_5 + 5\text{CO} \rightarrow \text{I}_2 + 5\text{CO}_2$ a 80.0 grams
 of iodine v oxide I_2O_5 reacts with 28.0

grams of carbon monoxide co
determine the mass of iodine i2 which
could be produced 80 g i2o5 1 mol i2o5
1 mol i2 xs 1 333 8 g i2o5 1 mol i2o5 28
g co 1 mol co

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reaction is the ratio of the actual yield
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