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ΔH must also be halved 3 answer c explanation option c is incorrect because 566 kJ/mol is the ΔH per mole of O_2 and not for 2 moles

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web we will revisit this today and go on to see that if we use a special kind of thermochemical reaction called the standard enthalpy of formation we can calculate enthalpies of reactions without having to manipulate a series of individual thermochemical equations for each step

5 s thermochemistry summary

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web we can express the first law in many ways one of the more useful expressions is that the change in internal energy E of a system in any process is equal to the heat q added to the system plus the work w done on the system by its surroundings E

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