

Fenomenos De Transporte

Camilla Rothe

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 density of the solution at that temperature $m = \frac{M}{d}$
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 solution of hcl is 1.0585 g/ml the molality of the solution is 1
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 of solution when studying the properties of solutions related
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 which can be abbreviated as molar or text m m pronounced
 molar the molar concentration of the solute is sometimes
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 of solution also known as molar concentration a capital m
 signifies solutions labelled with molar concentration a 1.0 m
 solution contains 1 mole of solute per litre of solution
 molality is the number of solvent moles per kilogram
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web 1 molarity mv mass molar mass $x = \frac{0.4500 \text{ L} \times 0.825 \text{ g/L}}{9579 \text{ g/mol} \times 0.0129 \text{ m}^2}$ molality $\frac{0.825 \text{ g/L} \times 141.9579 \text{ g/mol}}{0.00581158 \text{ mol} \times 0.00581158 \text{ mol} \times 0.4500 \text{ kg} \times 0.0129 \text{ m}^3}$ mole
 fraction $\frac{n_2}{n_1 + n_2} = \frac{0.4500 \text{ g} \times 18.015 \text{ g/mol}}{0.4500 \text{ g} \times 18.015 \text{ g/mol} + 0.00581158 \text{ mol} \times 45.00 \text{ g/mol}}$

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 4 2 2 a first convert 250 ml to liters $\frac{250}{1000} = 0.25$ then
 calculate molarity 5 moles 0.25 liters 20 m 3 c a solution
 with molarity 2 requires 2 m of n a oh per liter so $4 \times 2.8 \text{ m}^4$
 a a solution of molarity 1.5 m requires 1.5 mol of na to every
 litre of solvent

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of a mixture mmix can be calculated using the following formula $m_{\text{mix}} = \frac{m_1}{V_1} + \frac{m_2}{V_2} + \frac{m_3}{V_3} + \dots$ 480 520 720 624 1000 1 344 m q5 with increase in temperature which of these changes a molality b fraction of solute present in water

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