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Sabine Zange

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 reactant that is left over is described as being in excess

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 what is a limiting reactant it s not immediately obvious when
 you look at a chemical equation for a reaction when you see
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 are completely used up and all made into product

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 limits the amount of product s that can be obtained is the
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 the number of moles of each reactant present and compare
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 especially a directly reacting substance present at the
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amounts of reactants and products are typically measured by
 mass or by volume it is first necessary to convert the given
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 concentration in a chemical reaction its amount determines
 the amount of product made from that reaction see also
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 totally consumed it limits the reaction from continuing
 because there is none left to react with the in excess
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 grams of zns will be formed c how many grams of the excess
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 mgo is formed 4

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 the reaction stoichiometry thus limiting the amount of
 product generated molecular equation chemical equation in
 which all reactants and products are represented as neutral
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