

# Kings Of War 3rd Edition Uncharted Empires Pdf

Michal Rosen-Zvi

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3 oh 4 5 32 0 14 mole moles of h 2 o 40  
18 2 2 moles therefore according to the  
equation

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chem 1 to describe gas mixtures where  
the partial pressure of a gas was its  
mole fraction time the total pressure p  
a x ap t mole fraction a left frac mole a  
moles total right sum n 1 all x n 1 mole  
fraction of substance a is often denoted  
by x a and

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x b 1 2 the multiplication of the mole  
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advantages of molar fraction mole  
fraction is not dependent on the  
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and xb mol b mol a mol b note in the  
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defined as the number of molecules of  
component a divided by the total  
number of molecules of all the  
components in the solution

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100 g mass of benzene 30 g therefore  
mass of carbon tet 100 30 g benzene s  
molar mass c6h6 6 12 6 1 g mol 1 78 g  
mol 1 hence c6h6 s number of moles 30  
78 mol 0 3846 mol molar mass of

carbon tet ccl4 1 12 4 355 154 g mol 1  
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solution is 1 mol kg 1 or 1 m example  
02 11 11 g of urea nh2conh2 was  
dissolved in 100 g of water calculate  
the molarity and molality of the solution  
given n 14 h 1 c 12 o 16 given mass of  
solute urea 11 11 g mass of solvent  
water 100 g 0 1 kg to find molarity of  
solution solution

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is the mass of a sample of a 0 449 molal  
kbr that contains 2 92 kg of water 1  
molality moles solute divided by



kilograms solute 0.449 mol kg x 2.92 kg  
 x 1.31108 mol of kbr 2 moles times  
 molar mass equals grams 1.31108 mol  
 119.0023 g mol 156 g kbr

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web notice that the mole fraction has no units on it and is written as a decimal value do not change it to percent note of caution you could see the term mole percent it is simply the mole fraction multiplied by 100 for example in the problem just below the mole fraction of cinnamic acid is 0.2885 its mole percent would be 28.85

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advantages over some of the other units of concentration unlike molarity mole fraction is

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