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Annelies Wilder-Smith

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101 964 g Al_2O_3 302 3 g Al_2O_3 26 982 g
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 following reaction $2\text{H}_2 + \text{CO} \rightarrow \text{CH}_3\text{OH}$
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 if 68 5 g of co is reacted with 8 6 g of
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 reaction $\text{NH}_4\text{NO}_3 + \text{Na}_3\text{PO}_4 \rightarrow \text{NH}_4\text{PO}_4 + \text{NaNO}_3$
 which reactant is limiting
 assuming we started with 30 0 grams of
 ammonium nitrate and 50 0 grams of
 sodium phosphate what is the

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 $\text{I}_2\text{O}_5 + \text{CO} \rightarrow \text{CO}_2 + \text{I}_2$
 80 0 grams
 of iodine v oxide I_2O_5 reacts with 28 0
 grams of carbon monoxide co
 determine the mass of iodine I_2 which
 could be produced 80 g I_2O_5 1 mol I_2O_5
 1 mol I_2 xs 1 333 8 g I_2O_5 1 mol I_2O_5 28
 g co 1 mol co

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