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**Arturo Cuomo**

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 $2 \text{ CO}_2(\text{g}) \rightarrow 2 \text{ C}(\text{s}) + 2 \text{ O}_2(\text{g}) \quad \Delta h = 2 \times 393.5 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) \rightarrow 2 \text{ C}(\text{s}) + 3 \text{ H}_2(\text{g}) \quad \Delta h = 84.7 \text{ kJ}$   
 $\text{H}_2\text{O}(\text{g}) \rightarrow \text{H}_2(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) \quad \Delta h = 241.8 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$  from the following data  
 $\text{C}_2\text{H}_4(\text{g}) + 3 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 2 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = -1411 \text{ kJ}$   
 $\text{C}_2\text{H}_6(\text{g}) + 7 \text{ O}_2(\text{g}) \rightarrow 2 \text{ CO}_2(\text{g}) + 3 \text{ H}_2\text{O}(\text{l}) \quad \Delta h = -1560 \text{ kJ}$   
 $\text{H}_2(\text{g}) + \frac{1}{2} \text{ O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{l}) \quad \Delta h = -285 \text{ kJ}$

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 $\text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g}) \rightarrow \text{C}_2\text{H}_6(\text{g})$  given the following information

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cohen 2016 the equations for the reactions in the experiment done are as follows  $\text{NaOH(s)} + \text{Na(aq)} \rightarrow \text{OH(aq)}$

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ability to do work or transfer heat work is the transfer of energy from one body to another in a sense work is energy in the process of transfer

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web thus for reaction 9 2 in which 1 mole of  $\text{H}_2\text{O}$  and 1 mole of  $\text{OH}^-$  combine to form 2 moles of  $\text{H}_2\text{O}$   $\Delta H = -55.8 \text{ kJ}$  as shown note that the enthalpy change  $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$  is positive if heat is absorbed that is if  $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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the study of energy exchanged between the system and the surroundings

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