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

















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amu 1 12 of one carbon atom

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balanced molecular equations for the chemical changes ii write down the number of moles below the formula of each of the reactants and products iii write down the relative masses of the reactants and products with the help of formula below the respective formula these shall be the theoretical amounts of reactants and

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for jars d e f use the number of

molecules in the circle to determine the
number of moles of O_2 in the jar then
calculate the mass of the gas in the jar
d e f 12 mol O_2 2 mol O_2 6 mol O_2

mass 384 g mass 64 g mass 192 g 3 jars
g h and i contain ammonia gas NH_3

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balanced equation it is known that 1
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2 2 14 0 g 28 0 g 1 mol of NH_3 is 14 0 g
3 1 0 g 17 0 g

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6 941 amu x 6 0151 amu 1 x 7 0160
amu 6 941 1 0009x 7 0160 1 0009x 0
075 x 0 075 x 0 075 corresponds to a
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or 92 5 percent 3 7 the unit factor
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01 g 3 16 00 g 73 89 g b molar mass
 CS_2 12 01 g 2 32 07 g 76 15 g c molar
mass $CHCl_3$ 12 01 g 1 008 g 3 35 45 g
119 37 g d molar mass $C_6H_8O_6$ 6 12 01
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