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M Lipman

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 3 $\text{C}_2\text{H}_6(\text{g}) \rightarrow \text{C}_2\text{H}_4(\text{g}) + \text{H}_2(\text{g})$ $\Delta H = 94.5 \text{ kJ}$
 4 $\text{H}_2(\text{g}) + \frac{1}{2}\text{O}_2(\text{g}) \rightarrow \text{H}_2\text{O}(\text{g})$ $\Delta H = 71.2 \text{ kJ}$
 5 $\text{C}_2\text{H}_6(\text{g}) + \frac{7}{2}\text{O}_2(\text{g}) \rightarrow 2 \text{CO}_2(\text{g}) + 3 \text{H}_2\text{O}(\text{g})$ $\Delta H = 283 \text{ kJ}$

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web thus for reaction 9 2 in which 1 mole of H_3O^+ and 1 mole of OH^- combine to form 2 moles of H_2O $\Delta H = -55.8 \text{ kJ}$ as shown note that the enthalpy change $\Delta H_{\text{products}} - \Delta H_{\text{reactants}}$ is positive if heat is absorbed that is if $\Delta H_{\text{products}} > \Delta H_{\text{reactants}}$

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web hess s law enthalpy of reaction and hess s law abstract the goal of this lab is to perform experiments and practice hess s law by determining enthalpy change in chemical reactions in this lab three experiments were conducted in order to test hess s law the experiments required observing and collecting data on three different reactions

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web objective the purpose of the lab is to combine two reactants in the calorimeter and measure the heats of reaction in solution eventually allowing an individual to prove or disprove hess s law introduction thermodynamics is the study of energy exchanged between the system and the surroundings

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web 3 12 hess s law in this experiment you will measure the amount of heat released in these three related exothermic reactions $\text{NaOH}(\text{s}) + \text{Na}(\text{aq}) + \text{OH}(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ $2\text{NaOH}(\text{s}) + \text{H}(\text{aq}) + \text{Cl}(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$ $\text{Na}(\text{aq}) + \text{Cl}(\text{aq}) \rightarrow \text{H}_2\text{O}(\text{l})$