

Ensayo Sobre El Reciclaje

G Orfield

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 mixture of two substances $c_e a$ and $c_e b$ the mole fractions of

each would be written as

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 solvent water n_a 54 g 18 g mol 1 3 mol total number of
 moles $n_a + n_b$ 0 5 3 3 5 mol mole fraction of solute ethyl
 alcohol x_b $n_b / (n_a + n_b)$ 0 5 3 5 0 1429

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 and water in a solution prepared by dissolving 4 5 g of
 alcohol in 40 g of h_2o molar mass of h_2o is 18gm mole and
 molar mass of ch_3oh is 32gm mole solution moles of ch_3oh
 $4 5 / 32 = 0 14$ mole moles of h_2o $40 / 18 = 2 2$ moles therefore
 according to the equation

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 mixtures where the partial pressure of a gas was its mole
 fraction time the total pressure $p_a = x_a p_t$ mole fraction a left
 frac mole a moles total right sum n_1 all x_n 1 mole fraction
 of substance a is often denoted by x_a and
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 is equal to one $x_a + x_b = 1$ 2 the multiplication of the mole
 fraction by 100 gives the mole percentage 3 mole fraction is
 a unitless and dimensionless expression advantages of molar
 fraction mole fraction is not dependent on the temperature

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web since the mole fraction of the O_2 is 0.050 it contributes 0.050 or if you will 5.0 of the total pressure thus total pressure is 0.21 0.050 4.2 atm problem 9 a sample of 1.43 g of helium and an unweighed quantity of O_2 are mixed in a flask at room temperature

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