

Sartre Age Of Reason

SJ Ball

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this video accompanies the chm 252 lab nitration of bromobenzene in this video students will get an introduction to the lab as well as the set up for conducting the reaction students should

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for the nitration 4.01 ml of nitric acid 4.00 ml of sulfuric acid and 4.56 ml bromobenzene were used to complete the reaction with these measurements bromobenzene was found to be the limiting reagent since only 4.56 ml of bromobenzene was used the amount of product predicted to be recovered was 8.7533 g

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the aim of the experiment is to determine and study electrophilic aromatic substitution reaction especially nitration of methyl benzoate to give methyl m nitrobenzoate thus determine the weight melting point and percentage yield of the pure methyl m nitrobenzoate abstract

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in this experiment we will carry out the nitration of bromobenzene bromine substituted on an aromatic ring is an ortho para directing group it will direct an incoming electrophile to the ortho and para positions it is also deactivating meaning that its reactions will be slower than the identical reactions of benzene

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chemical reaction of nitration consists of a nitro group being added to or substituted in a molecule nitration can basically be carried out by a mixture of concentrated nitric acid and sulfuric acid this mixture is useful to obtain the active nitronium ion

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1 go to the restroom and wash the affected area with lots of soap and water the soap in the lab is not suitable for this purpose 2 return to the lab and apply mineral oil to the affected area 3 the summary to this warning is not to touch any part of your body while performing this experiment if you must leave the lab for any reason 1

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in the nitration of bromobenzene a mixture of sulfuric acid and nitric acid react with one another nitric acid the weaker of the two acids will serve as base to form nitronium ion bromobenzene will attack the nitronium ion by nucleophilic attack
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experiment was performed in order to
demonstrate the principles of nitration
through electrophilic aromatic
substitution it also examined the effects
of bromine on an aromatic ring ortho
para vs meso directing

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then actual is your actual findings yield
actual theoretical x 100 what are the
possible side reactions of nitration of
bromobenzene dinitration can occur at
ortho para or at both ortho positions
occurs in overheating meta product
happens very rarely because bromine is
an ortho para director

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the soap in the lab is not suitable for
this purpose 2 return to the lab and
apply mineral oil to the affected area
theory and explanation of the nitration
of bromobenzene if the temperature of
the reaction mixture exceeds 60 °C
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the bromine substituent in
bromobenzene is an ortho para director
the para position is favored due to
steric effects in the reaction that takes
place in this lab a nitro group is
substituted onto bromobenzene at the
para position to form 1-bromo-4-
nitrobenzene

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study with quizlet and memorize
flashcards containing terms like
reacting sulfuric and nitric acid
together produces the nitronium ion
alone nitric acid is not a strong enough
electrophile to react with
bromobenzene sulfuric acid is added to
the nitric acid the reverse rxn can
overheat and boil easily includes a
claisen adaptor and

nitration of bromobenzene 6
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over 5 minutes slowly pipet this
bromobenzene into the icy cold
nitration mixture HNO_3 3 mL H_2SO_4 4 mL in the
125 mL flask and stir continue to stir
the reaction mixture in the ice bath for
30 minutes do not heat in a new 250 mL
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on the right the arenium cation has a
nitro set with incredibly small 0.6
Å bond distances this will result
in an optimization failure as the atoms
are even close for the programme to
accommodate their electrons
demonstration of creating the p
arenium cation by nitration of
bromobenzene

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experiment the electrophilic aromatic
substitution mechanism was used to
obtain 4-nitrobromobenzene and 2-
nitrobromobenzene recrystallization
was used to purify 4

nitrobromobenzene

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the purpose of this lab was to observe where substituents are added on a benzene ring by way of electrophilic aromatic substitution the electrophilic source in the nitration of bromobenzene comes from a nitronium ion NO_2^+ which acts as the electrophilic species

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nitration of bromobenzene is an electrophilic aromatic substitution reaction the electrophile will be formed from 4 0ml of nitric acid and 4 0ml of sulfuric acid and will then react with 4 5ml of bromobenzene this will produce o bromonitrobenzene and p bromonitrobenzene

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- Feb 22 2023 lab report 10 nitration of bromobenzene raekwon filmore cm 244 section 40 march 27 2018 introduction for this experiment nitration of bromobenzene was the focus of the lab the benzene is an aromatic compound and when it reacts with wither a mixture of sulfuric acid or nitric acid

creates what is known as a nitro group **oxford new english file intermediate cevapları erkan rua** - Jun 14 2023

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