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Chao Zhang

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web solution $v_1 = 2.50 \text{ l}$ $v_2 = ?$ $t_1 = 273 \text{ K}$ $t_2 = 273 + 70 = 343 \text{ K}$ we will be applying charles law as pressure is constant here $v_1/t_1 = v_2/t_2$ or $v_2 = v_1 \cdot t_2/t_1 = 2.50 \text{ l} \cdot 343 \text{ K} / 273 \text{ K} = 3.12 \text{ l}$ at 90 c a helium sample has a volume of 500 ml determine the temperature at which the volume of the liquid will become 240 ml