

# Design Failure Mode Effect Analysis

**Emilie Sanchez**

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342g of 20 by mass of  $\text{Ba(OH)}_2$  solution sp gr 0.57 is reacted with 1200 ml of 2M  $\text{HNO}_3$  if the final density of solution is same as pure water then molarity of the iron in resulting solution which decides the nature of the above solution is 1.025M, 2.05M, 3.088M, 4. none of

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web 1 molarity  $M = \frac{m}{V}$  mass molar mass  $\times 0.4500 \text{ L} = 0.825 \text{ g/L}$  9579 g/mol  $\times 0.0129 \text{ M} = 2 \text{ molality}$  0.825 g/L 9579 g/mol 0.00581158 mol 0.00581158 mol 0.4500 kg 0.0129 M 3 mole fraction  $n_2 = \frac{4.0825 \text{ g/L}}{4.0825 \text{ g/L} + 9579 \text{ g/mol} \times 0.00581158 \text{ mol}} = 0.4500$  0 g/L 0.015 g/mol 24 97918401 mol

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litre of solvent

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web answer the correct option is c explanation the molarity of a mixture  $m_{mix}$  can be calculated using the following formula  $m_{mix} = \frac{m_1}{V_1} + \frac{m_2}{V_2} + \frac{m_3}{V_3} + \dots$   $\frac{480}{1000} + \frac{520}{1000} = 1.000$   $\frac{1000}{1000} = 1$   $\frac{344}{1000} = 0.344$   $m_{q5}$  with increase in temperature which of these changes a molality b fraction of solute present in water

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