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Mark Steyvers

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the reaction below given the following reactions and
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 $\text{C}_2\text{H}_6 \text{ g} \rightarrow \text{C}_2\text{H}_4 \text{ g} + \text{H}_2 \text{ g}$ $\Delta h = 137 \text{ kJ}$ $\text{C}_2\text{H}_6 \text{ g} \rightarrow \text{C}_2\text{H}_2 \text{ g} + 2 \text{H}_2 \text{ g}$ $\Delta h = 283 \text{ kJ}$

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reaction $\text{C}_2\text{H}_4 \text{ g} + \text{H}_2 \text{ g} \rightarrow \text{C}_2\text{H}_6 \text{ g}$ from the following data $\text{C}_2\text{H}_4 \text{ g} + 3 \text{O}_2 \text{ g} \rightarrow 2 \text{CO}_2 \text{ g} + 2 \text{H}_2\text{O} \text{ l}$ $\Delta h = 1411 \text{ kJ}$ mole $\text{C}_2\text{H}_6 \text{ g} + 2 \text{O}_2 \text{ g} \rightarrow 2 \text{CO}_2 \text{ g} + 3 \text{H}_2\text{O} \text{ l}$ $\Delta h = 1560 \text{ kJ}$

2 g 2 co 2 g 3 h 2 o l h 1560 kj mole h 2 g 1 2 o 2 g h 2 o l h 285 kj mole

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