

The Science Of Cooking

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species that increases the concentration of hydroxide ions in aqueous solution an example of an arrhenius base is the highly soluble sodium hydroxide NaOH sodium hydroxide dissociates in water as follows $\text{NaOH} \rightarrow \text{Na}^+ + \text{OH}^-$ in water sodium hydroxide

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web the arrhenius theory of acids and bases acids are substances that produce hydrogen ions in solution bases are substances that produce hydroxide ions in solution neutralization happens because hydrogen ions and hydroxide ions react to produce water $\text{H}^+ + \text{OH}^- \rightarrow \text{H}_2\text{O}$ $\text{HCl} + \text{H}_2\text{O} \rightarrow \text{H}_3\text{O}^+ + \text{Cl}^-$ hydrochloric

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the h 3 o the reversible reaction contains two acids and two
bases we think of them in pairs called conjugate pairs when
the acid ha loses a proton it forms a base a when the base a
accepts a proton back again it obviously reforms the acid ha

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molecular unit contains at least one

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predict a wider variety of acid base reactions lewis theory
used electrons instead of proton transfer and specifically
stated that an acid is a species that accepts

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learning chemistry across different contexts one such
foundational principle is the acid base behavior of molecules
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is stressed because it lends itself well to studying equilibrium
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