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we consider using titration to measure the amount of aspirin in a solution using titration it would be difficult to identify the end point because aspirin is a weak acid and reactions may proceed slowly using back titration the end point is more easily recognised in this reaction as it is a reaction between a strong base and a strong acid

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back titration is used for 1 volatile substances e.g.  $\text{NH}_3$  2 insoluble or slightly soluble substances e.g.  $\text{CaCO}_3$  3

substances for which the quantitative reaction proceeds rapidly only in the presence of excess of reagent e g lactic acid aspirin 4 substances which decompose on heating e g

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1 acid base salt

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and it

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cm<sup>3</sup> of 1 00 mol l<sup>-1</sup> sodium hydroxide solution

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