

Appropriation In Art Meaning

N Colangelo

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object made up of copper which is put
in the calorimeter are the same the
initial temperature of the calorimeter
and water is 30 c and that of hot object
is 60 c the specific heat of copper and
water are 0 09 cal gm c and 1 cal gm c
respectively

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equation $q = \text{specific heat} \times m \times \Delta t$ where
 q is heat flow m is mass in grams and
 Δt is the temperature change plugging
in the values given in the problem you
get $q_{\text{water}} = 418 \text{ J/g} \times 110 \text{ g} \times 26.6 \text{ }^\circ\text{C}$
 $= 250 \text{ J}$ $q_{\text{water}} = 550 \text{ J}$ $\Delta h = q_{\text{water}} = 550 \text{ J}$

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207 J of heat is added 2

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the heart of all calorimetry problems
and calculations because the heat
released or absorbed at constant
pressure is equal to Δh the relationship
between heat and Δh_{rxn} is $\Delta h_{\text{rxn}} = q_{\text{rxn}}$
 $q_{\text{calorimeter}} = m c \Delta t$ 7 3 14 7 3 14 Δh_{rxn}
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1450 c what will the change in enthalpy
be 2 sketch the heating curve
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1 3 i have 18.7 grams of an unknown
liquid chemical at a temperature of 350
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the solutions are 1 00 g ml and that
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