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 displacement 2  $\text{MgCl}_2 + \text{Li}_2\text{CO}_3 \rightarrow \text{MgCO}_3 + 2\text{LiCl}$  double  
 displacement 3  $\text{C}_6\text{H}_{12} + 9\text{O}_2 \rightarrow 6\text{CO}_2 + 6\text{H}_2\text{O}$  combustion 4  $\text{Pb} + \text{FeSO}_4 \rightarrow \text{PbSO}_4 + \text{Fe}$  single displacement 5  $\text{CaCO}_3 \rightarrow \text{CaO} + \text{CO}_2$   
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web 1  $3\text{NaBr} + \text{H}_3\text{PO}_4 \rightarrow \text{Na}_3\text{PO}_4 + 3\text{HBr}$  type of reaction double  
 displacement 2  $3\text{Ca(OH)}_2 + \text{Al}_2\text{SO}_4 \rightarrow 3\text{CaSO}_4 + 2\text{Al(OH)}_3$  type of  
 reaction double displacement 3  $3\text{Mg} + \text{Fe}_2\text{O}_3 \rightarrow 2\text{Fe} + 3\text{MgO}$  type of  
 reaction single displacement 4  $\text{C}_2\text{H}_4 + 3\text{O}_2 \rightarrow 2\text{CO}_2 + 2\text{H}_2\text{O}$  type of  
 reaction combustion 5  $2\text{PbSO}_4 + 2\text{PbSO}_3 + \text{O}_2 \rightarrow 2\text{PbO}_2 + 2\text{PbSO}_4$  type of reaction

**chemical reactions types definitions and examples chemistry** - Feb 25 2022

web a chemical reaction is a process in which one or more  
 substances are converted to one or more different  
 substances the starting substances are called the reactants  
 and the new substances that form are called the products

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web study guide chemical reactions 1 give an example of a  
 chemical reaction leaves turning color fireworks food  
 spoiling 2 differentiate between a physical and chemical

change physical does not change the identity of the matter

can be changed back easily chemical new substance is  
created with new physical and chemical