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chemistry is the transfer of protons
thus logic suggests a relation between
the relative strengths of conjugate acid
base pairs the strength of an acid or
base is quantified in its ionization
constant K_a or K_b which represents the
extent of the acid or base ionization
reaction

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 $K_a = 1.8 \times 10^{-5}$ and acetate is a weak
base $K_b = 5.6 \times 10^{-10}$ like acids
strong and weak bases are classified by
the extent of their ionization strong
bases disassociate almost or entirely to
completion in aqueous solution

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acids bases and enzymes many acids and bases in living things provide the ph that enzymes need enzymes are biological catalysts that must work effectively for biochemical reactions to occur most enzymes can do their job only at a certain level of acidity cells secrete acids and bases to maintain the proper ph for enzymes to do their work acids and bases definition examples

properties uses with - May 08 2022 acid base titration difference between acid and base acids and bases definition an acid is any hydrogen containing substance that is capable of donating a proton hydrogen ion to another substance a base is a molecule or ion able to accept a hydrogen ion from an acid acidic substances are usually identified by their sour taste 5 1 acids nomenclature chemistry

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electricity because they contain dissolved ions therefore acids and bases are electrolytes strong acids and bases will be strong electrolytes weak acids and bases will be weak electrolytes

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arrhenius s definition of acids and bases the earliest definition of acids and bases is arrhenius s definition which states that an acid is a substance that forms hydrogen ions h when dissolved in water and a base is a substance that forms hydroxide ions oh when dissolved in water

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the reaction of an acid and a base is called a neutralization reaction although acids and bases have their

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